

SLIP TEST-2
CHAPTER-2 : CHEMICAL REACTIONS AND EQUATIONS

Name:..... Section:..... Roll No:..... Max.Marks:20

I. Answer the following questions. Each carries four marks. 2 x 4 = 8 M

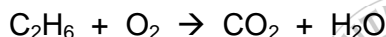
- 1) How many types of chemical reactions are there? Explain with examples.
- 2) Take two beakers and prepare lead nitrate aqueous solution and potassium iodide aqueous solutions. What are the colours of the solutions. Now mix them in another beaker. What happens? What type of chemical reaction it is? What are the products?

II. Answer the following questions briefly. Each carries two marks. 2 x 2 = 4 M

- 3) Give two examples for chemical reactions in which precipitate is formed?
- 4) If 40gm of methane is burnt, then how much amount of carbon dioxide is released?

III. Answer the following in one or two sentences. Each carries one marks. 2 x 1 = 2 M

- 5) How can we prevent rusting of iron?
- 6) Balance the following chemical equation.



IV. Choose the correct choice and write down in the given brackets. 6 x 1 = 6 M

- 7) The gas filled in potato chips flush bags []
 A. Hydrogen B. Oxygen C. Nitrogen D. Chlorine
- 8) Galvanising means, coating of On the iron substances to prevent corrosion []
 A. Zn B. Cr www.ignitephysics.net
 C. Cu D. C
- 9) $2\text{Fe}_2\text{O}_3 + 3\text{C} \rightarrow 4\text{Fe} + 3\text{CO}_2$ Then which is true of the following []
 A. Carbon is oxidised B. Carbon is reduced
 C. Iron is oxidised D. Iron oxide is oxidised
- 10) It converts slaked lime into milk white substance []
 A. Oxygen B. Carbon dioxide
 C. Hydrogen D. Sulphur dioxide
- 11) A substance is in light yellow colour. If we put it in sun light, it changes into gray colour. What is the substance? []
 A. Lead Iodide B. Potassium Iodide
 C. Silver Bromide D. Hydrogen Chloride
- 12) The following image relates to the reaction []
 A. $\text{CuSO}_4 + \text{Fe}$
 B. $\text{FeSO}_4 + \text{Cu}$
 C. $\text{CuSO}_4 + \text{Zn}$
 D. $\text{ZnSO}_4 + \text{Cu}$

