

FORMATIVE ASSESSMENT-3
CHAPTERS – 8,9,10

Name:..... Section:..... Roll No:..... Max.Marks:25

I. Answer the following questions. Each carries four marks. 2 x 4 = 8 M

- 1) Define ionization energy. What are the affecting factors of ionization energy? Explain.
2) Write a brief notes about Quantum numbers.

II. Answer the following questions briefly. Each carries two marks. 3 x 2 = 6 M

- 3) Draw Moeller's chart of showing ascending order of energies of various atomic orbitals.
4) Explain the formation of CaO molecule.
5) How the Electro negativity varies in a period and in a group in periodic table of elements.

III. Answer the following in one or two sentences. Each carries one marks. 3 x 1 = 3 M

- 6) How many maximum number of electrons can be accommodated in all d- orbitals in M-shell?
7) State Modern periodic law.

- 8) Name two molecules having double bond.

IV. Choose the correct choice and write down in the given brackets. 4 x 1 = 4 M

- 9) Identify the covalent compound from the following []

A. $MgCl_2$ B. $BeCl_2$ C. $NaCl$ D. $AlCl_3$

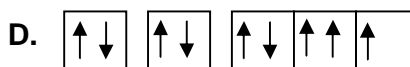
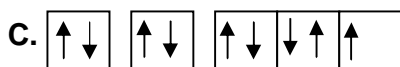
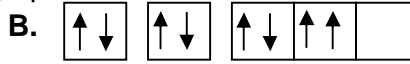
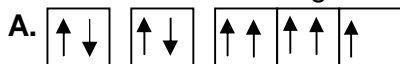
- 10) The element having greatest value of Electron affinity []

A. Fluorine B. Chlorine C. Lithium D. Sodium

- 11) Valence electronic configuration of *Chromium* is []

A. $4s^23d^4$ B. $4s^13d^5$ C. $4s^23d^9$ D. $4s^13d^{10}$

- 12) Correct method of filling electrons in 1s, 2s, 2p orbitals. []



IV. Fill in the blanks with suitable answers. 4 x 1 = 4 M

- 13) has the greatest value of Electro negativity.

- 14) The element having atomic number 48 belongs to block in periodic table.

- 15) proposed by ionic bond.

- 16) In nl^x method; indicates principal energy level.