

FORMATIVE ASSESSMENT-3
CHAPTERS – 8,9

Name:..... Section:..... Roll No:..... Max.Marks:25

I. Answer the following questions. Each carries four marks. 2 x 4 = 8 M

- 1) Define ionization energy. What are the affecting factors of ionization energy? Explain.
2) Write a brief notes about Quantum numbers.

II. Answer the following questions briefly. Each carries two marks. 3 x 2 = 6 M

- 3) Draw Moeller's chart of showing ascending order of energies of various atomic orbitals.
4) Write differences between an Orbit and an orbital.
5) How the Electro negativity varies in a period and in a group in periodic table of elements.

III. Answer the following in one or two sentences. Each carries one marks. 3 x 1 = 3 M

- 6) How many maximum number of electrons can be accommodated in all d- orbitals in M-shell?
7) State Modern periodic law.
8) Name any four metalloids.

IV. Choose the correct choice and write down in the given brackets. 4 x 1 = 4 M

- 9) 1 pico meter (pm) = m []
A. 10^{-12} B. 10^{-13} C. 10^{-14} D. 10^{-15}

- 10) The element having greatest value of Electron affinity []
A. Fluorine B. Chlorine C. Lithium D. Sodium

- 11) Valence electronic configuration of *Chromium* is []
A. $4s^23d^4$ B. $4s^13d^5$ C. $4s^23d^9$ D. $4s^13d^{10}$

- 12) Correct method of filling electrons in 1s, 2s, 2p orbitals. []
- A.

↑↓	↑↓	↑↑	↑↑	↑
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 B.

↑↓	↑↓	↑↓	↑↑	
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- C.

↑↓	↑↓	↑↓	↓↑	↑
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 D.

↑↓	↑↓	↑↓	↑↑	↑
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IV. Fill in the blanks with suitable answers. 4 x 1 = 4 M

- 13) has the greatest value of Electro negativity.
14) The element having atomic number 48 belongs to block in periodic table.
15) The value of 'l' for the sub stationary shell 'd' is
16) In nl^x method; indicates principal energy level.